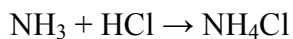
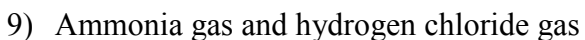
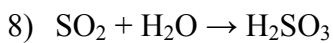
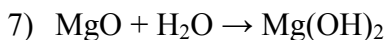
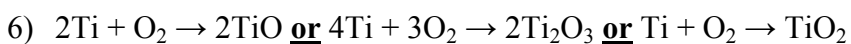
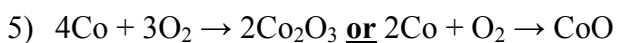
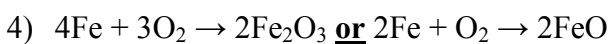
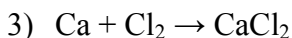
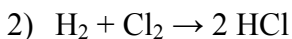


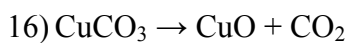
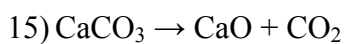
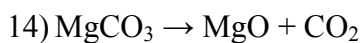
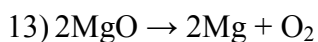
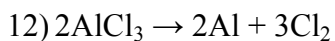
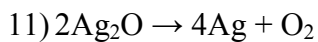
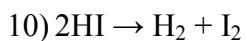
Reaction Prediction Answer Key

Write the balanced molecular equation along with [the net ionic equation](#).

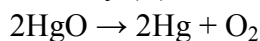
Synthesis Reactions



Decomposition

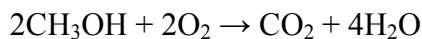


17) Mercury (II) oxide is a bright red powder. It decomposes on heating.



Combustion

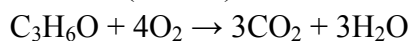
18) Methanol (CH_3OH)



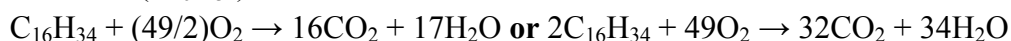
19) Octane (C_8H_{18})



20) Acetone ($\text{C}_3\text{H}_6\text{O}$)



21) Kerosene ($\text{C}_{16}\text{H}_{34}$)



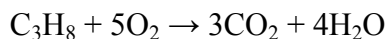
22) Butane (C_4H_{10})



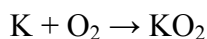
23) Ethanol (C_2H_6)



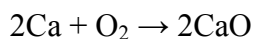
24) Propane (C_3H_8)



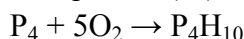
25) Potassium



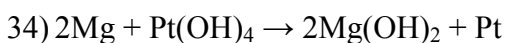
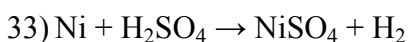
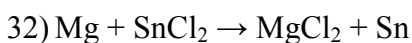
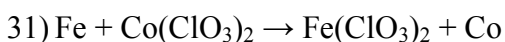
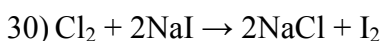
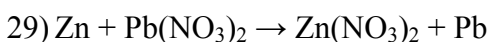
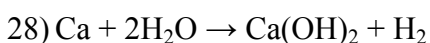
26) Calcium



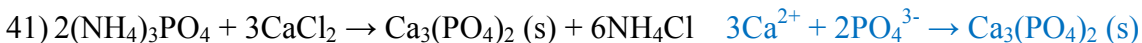
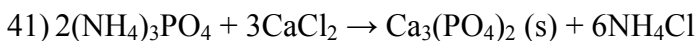
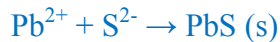
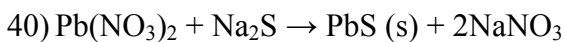
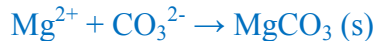
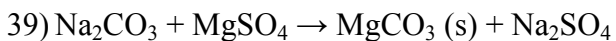
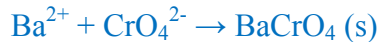
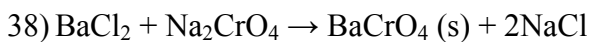
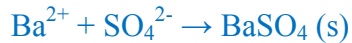
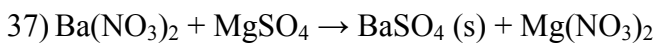
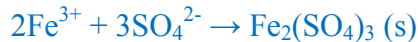
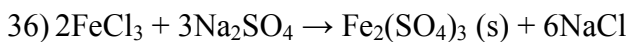
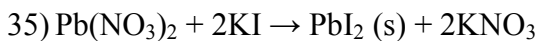
27) Phosphorus (P_4)



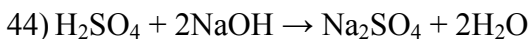
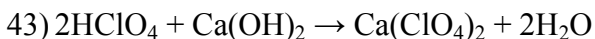
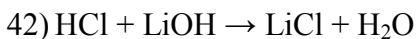
Oxidation-reduction



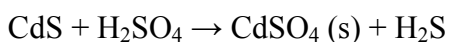
Double Displacement (Precipitation)



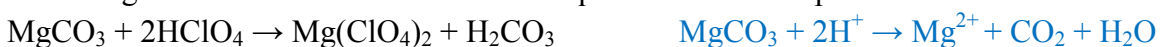
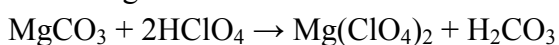
Double Displacement (Acid-Base)



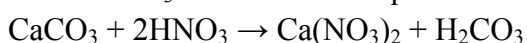
45) Solid cadmium sulfide reacts with an aqueous solution of sulfuric acid



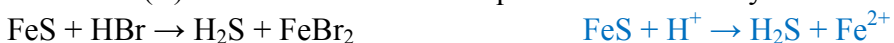
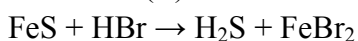
46) Solid magnesium carbonate reacts with an aqueous solution of perchloric acid*



47) Solid CaCO_3 reacts with an aqueous solution of nitric acid*



48) Solid iron (II) sulfide reacts with an aqueous solution of hydrobromic acid



* H_2CO_3 is very unstable so it becomes $\text{CO}_2 + \text{H}_2\text{O}$ very easily